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IIT-JEE2011-Code-0-Questions and Solutions-Paper-I and II 8 NARAYANA IIT ACADEMY $4 \times 1.6 \times 10^2 \times 10^{-10} \times 1.6 \times 10^{-14} \times 1.6 \times 10^{-17} = (1.6)^2 \times 10^{-14} \times 1.6 \times 10^{-7} \times 1.6 \times 10^{-17}$. The volume (in mL) of 0.1 M AgNO₃ required for complete precipitation of chloride ions present in 30 mL of 0.01 M solution of [Cr(H₂O)₅Cl]Cl₂, as silver chloride is close to Key. : 6

IIT JEE (2 011) PAPER II QUESTION & SOLUTIONS CODE 0

IIT JEE Advanced 2011 Question Paper with Answers : The Indian Institutes of Technology Joint Entrance Examination (IIT-JEE) was a yearly engineering entrance examination in India. It was used as the sole admission test by the fifteen Indian Institutes of Technology (IITs), Indian School of Mines and IIT BHU.

IIT JEE Advanced 2011 Question Paper With Answers ...

IIT-JEE2009-Code-0-Questions and Solutions-Paper-I and II 3 4. The Henry's law constant for the solubility of N₂ gas in water at 298 K is 1.0×10^5 atm. The mole fraction of N₂ in air is 0.8. The number of moles of N₂ from air dissolved in 10 moles of water at 298 K and 5 atm pressure is

IIT-JEE2009-Code-0-Questions and Solutions-Paper-I and II ...

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JEE(ADVANCED)-2018-Paper-2-PCM-3 FIITJEE Ltd., FIITJEE House, 29 -A, Kalu Sarai, Sarvapriya Vihar, New Delhi 110016, Ph 46106000, 26569493, Fax 26513942 website: www.fiitjee.com. Q.4 A wire is bent in the shape of a right angled triangle and is placed in front of a concave mirror of focal length f, as shown in the figure.

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AIEEE 2011 paper solutions

(2) Vidyalankar : IIT JEE 2011 Question Paper & Solution 2 Useful Data $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$ or $8.206 \times 10^2 \text{ L atm K}^{-1} \text{ mol}^{-1}$ $1 \text{ F} = 96500 \text{ C mol}^{-1}$ $h = 6.626 \times 10^{-34} \text{ Js}$ $1 \text{ eV} = 1.602 \times 10^{-19} \text{ J}$ $c = 3.0 \times 10^8 \text{ ms}^{-1}$ $N_A = 6.022 \times 10^{23}$

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